**Volumetric Titrations**

1. What volume of hydrochloric acid (concentration 0.1 mol l-1) is required to neutralise 100 cm3 of sodium hydroxide (concentration 0.5 mol l-1)?
2. If 25 cm3 of hydrochloric acid is neutralised by 50 cm3 of sodium hydroxide solution (concentration 2 mol l-1), what is the concentration of the acid?
3. What volume of potassium hydroxide (concentration 2 mol l-1) is required to neutralise 50 cm3 of sulphuric acid (concentration 1 mol l-1)?
4. If 50 cm3 potassium hydroxide is neutralised by 10 cm3 of sulphuric acid (concentration 2 mol l-1), what is the concentration of the alkali?
5. What volume of nitric acid (concentration 0.5 mol l-1) is required to neutralise 25 cm3 of sodium hydroxide solution (concentration 0.4 mol l-1)?
6. What volume of sulphuric acid (concentration 2.5 mol l-1) is required to neutralise 25 cm3 of potassium hydroxide solution (concentration 4 mol l-1)?